

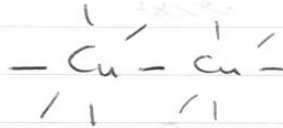
3	4	5	6	7	8	9	10	11	12
Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn
Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd
La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg

1 Yzer Ruthenium & Osmium

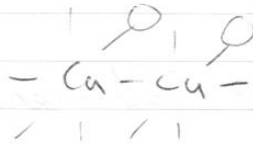
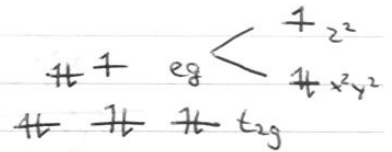


bind beter op S

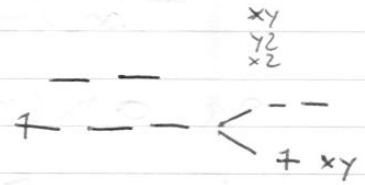
3



$\text{Cu}^{2+} \quad 11e \quad 2+ = 9e$

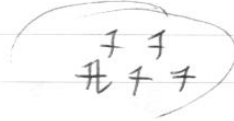


$\text{Ti} \quad 4e \quad 3+ \quad 1e$

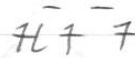


4 korte 2 lange

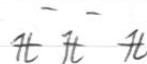
4  $[\text{CoF}_6]^{3-}$  weak  $9e \rightarrow 6$

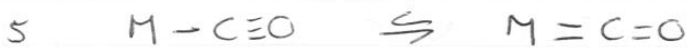


$[\text{Mn}(\text{CN})_6]^{3-}$  strong  $7e \rightarrow 4$



$[\text{IrCl}_6]^{3-}$  strong  $9e \rightarrow 6$

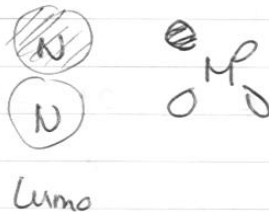




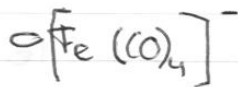
6



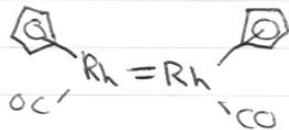
7



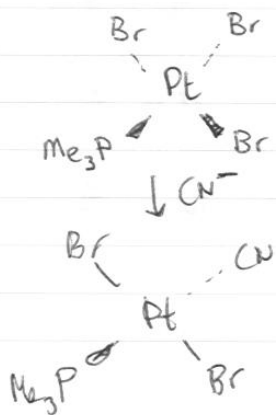
$\pi$  backbonding by  $\eta^2$  = sterker



b



9



10  
ab

1	oxidatieve additie	A	16e
2	ligand dissociatie	B	18e
3	ligand associatie	C	16e
4	ligand disso	D	18e
5	ligand associatie	E	16e
6	reductieve eliminatie	F	18e

9e

c oxidatie additie kost meer energie van het atoom dan bij H<sub>2</sub>